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Sep 16, 2010 Best Answer: Hydrogen bonding is a type of intermolecular force. When you compare it in strength to say, covalent bonding it is indeed very weak, being

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A hydrogen bond is a weak type of force that forms a special type of dipole-dipole attraction which occurs when a hydrogen atom bonded to a strongly electronegative

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hydrogen bond n. A chemical bond in which a hydrogen atom of one molecule is attracted to an electronegative atom, especially a nitrogen, oxygen, or fluorine atom

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The weak hydrogen bond was first identified in 1935, GAUTAM R. DESIRAJU and THOMAS STEINER. in International Union of Crystallography Monographs on

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A hydrogen bond is the electrostatic attraction between polar molecules that occurs when a hydrogen (H) atom bound to a highly electronegative atom such as nitrogen

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### **Hydrogen bonds - strong or weak? - chemical**

A quote from a Wiki page: "The hydrogen bond (5 to 30 kJ/mole) is stronger than a van der Waals interaction, but weaker than covalent or ionic bonds.

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a hydrogen bond is a type of weak chemical bond. It is formed between the slightly positively charged hydrogen atom of one molecule and a slightly negative charged

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The weak hydrogen bond, also known as non-conventional hydrogen bond, has been the subject of intense scrutiny over recent years in several fields, in particular

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Mar 14, 2012 Best Answer: It's weak because hydrogen gives up electrons easily. It only has 1 and there for wants to bond with other elements, although when it does, it

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Intermolecular hydrogen bonding is responsible for the high boiling point of water Intermolecular ("weak") van der Waals forces: London dispersion force; Hydrogen

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Hydrogen bonds are not really bonds but intermolecular forces - weak forces which arise between molecules. Hydrogen bonds are the strongest kind of intermolecular

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